Supporting Information

Functionalized Dibenzo[*g*,*p*]chrysenes: Variable Photo- and Electronic Properties and Liquid Crystal Crystal Chemistry

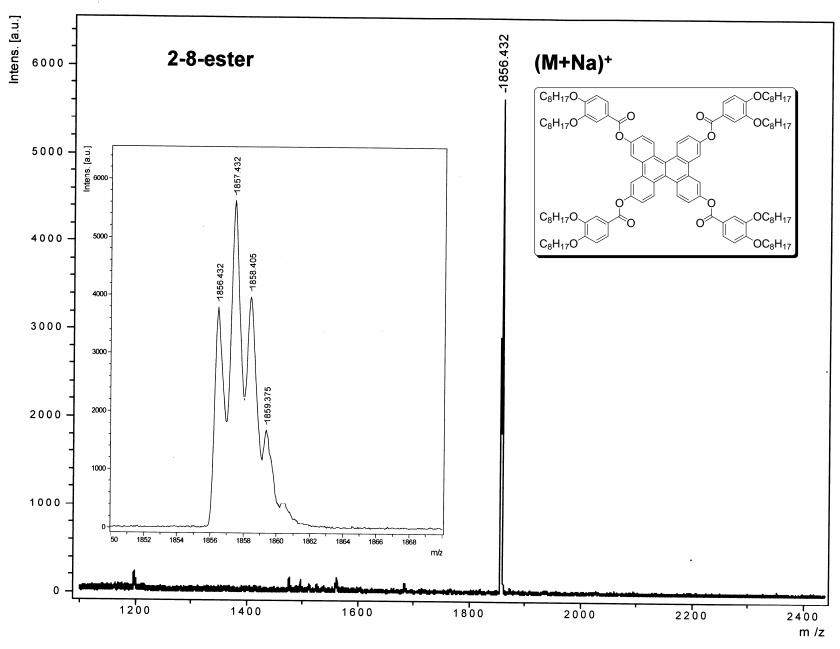
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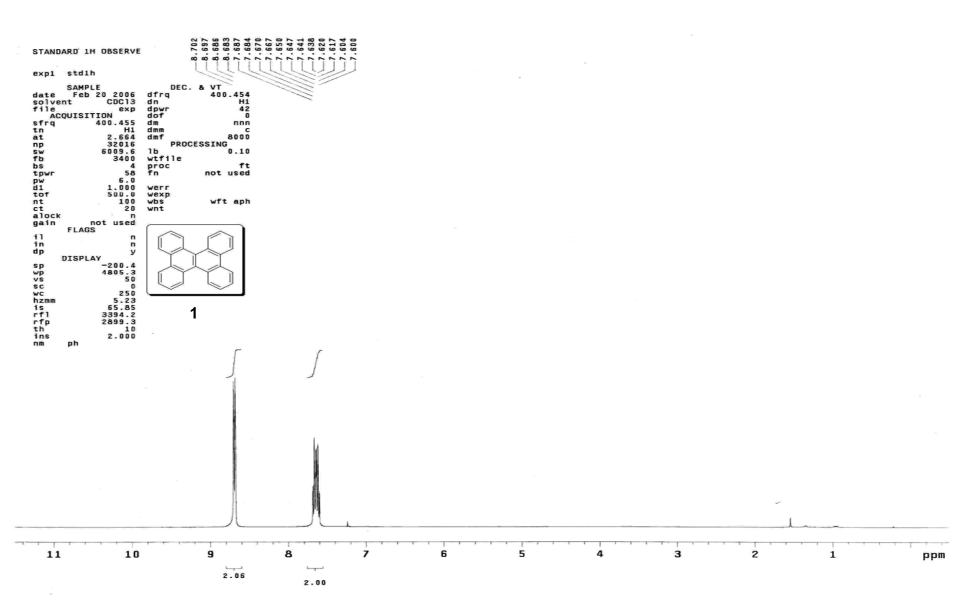
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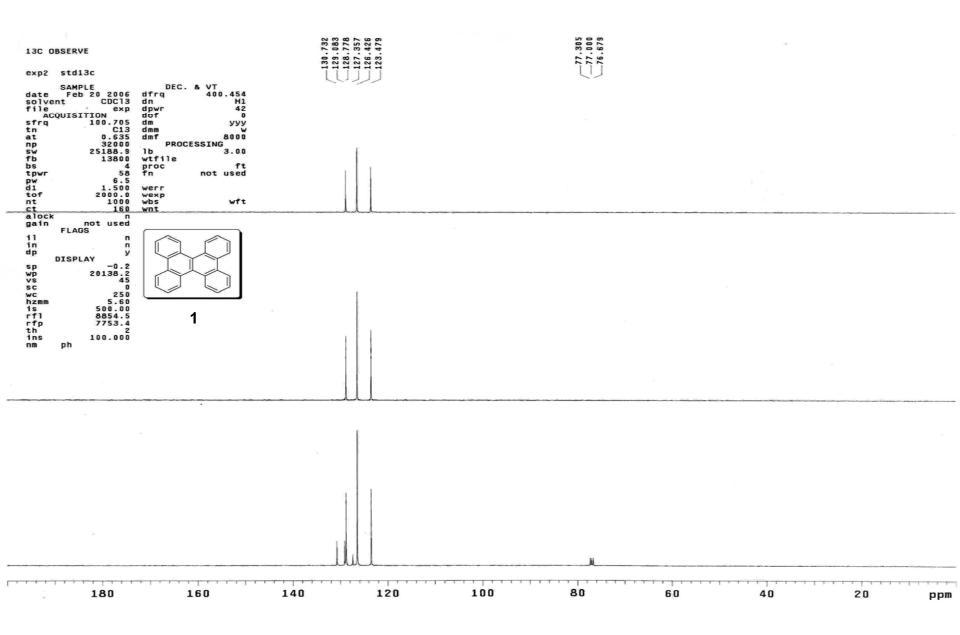
rsliu@mx.nthu.edu.tw

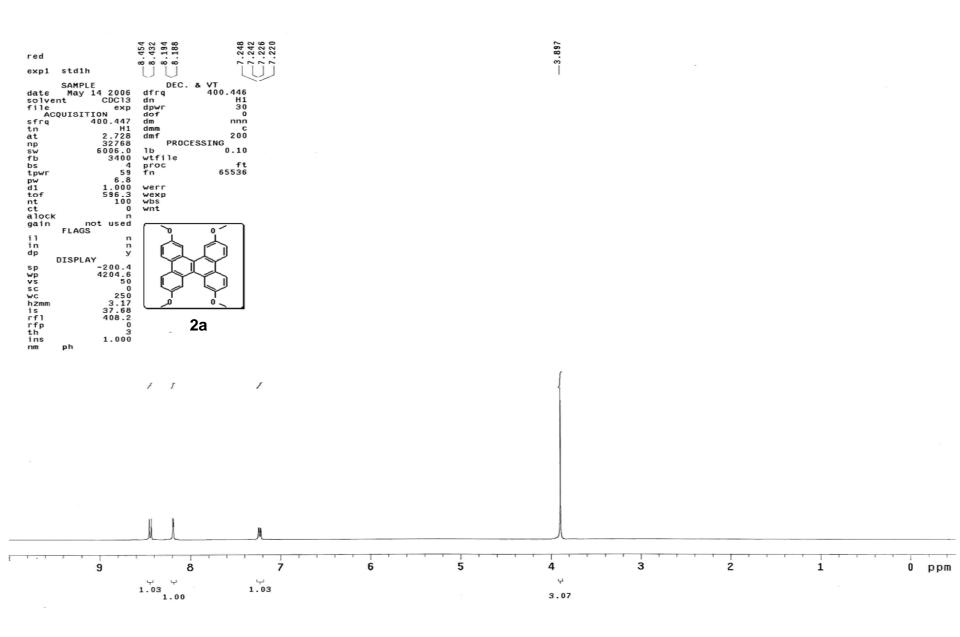
Contents:

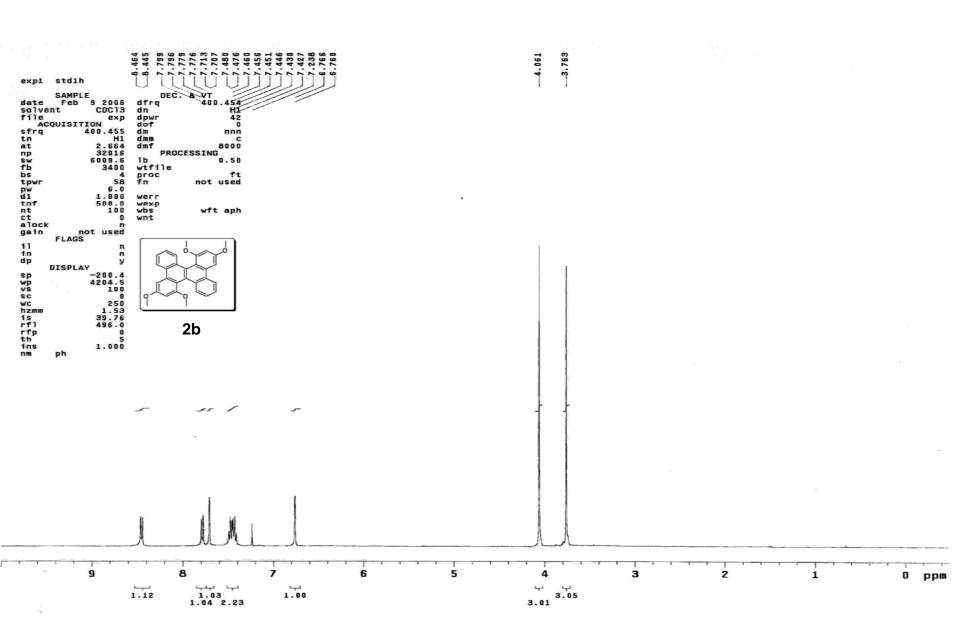
(1) MALDI-TOF (M+Na) ⁺ for compound 4	-S2
(2) NMR Spectra for key compounds	-S3
(3) Cyclic Voltammograms for key compound	S36
(4) The TGA measurement for 4	S45
S1	

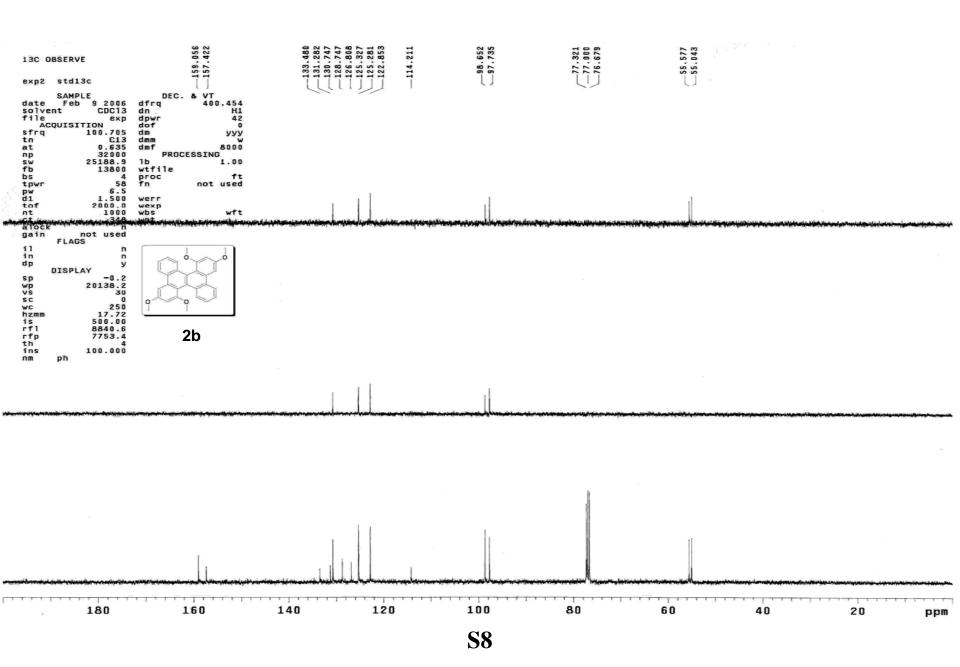


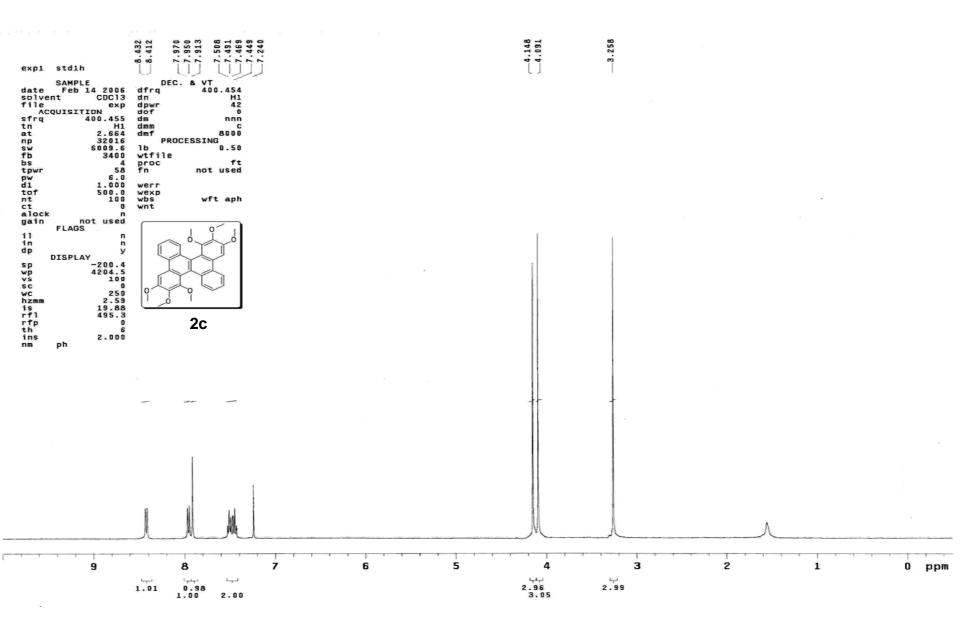


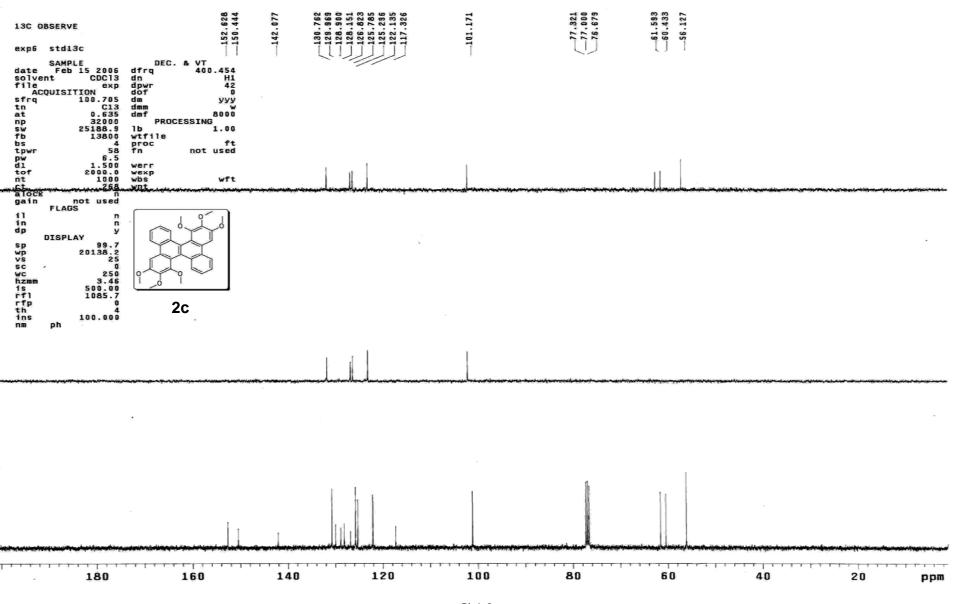


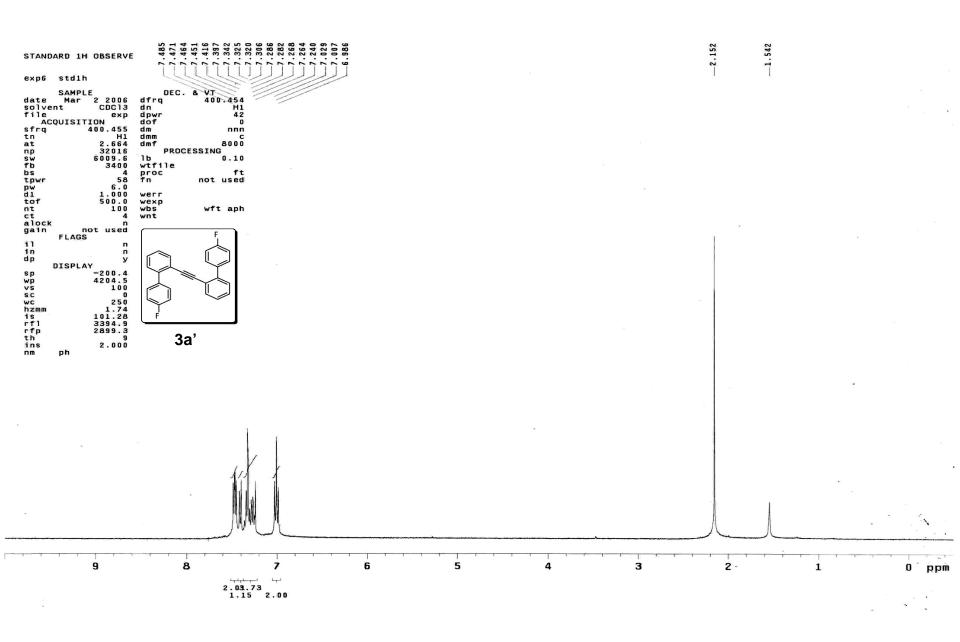


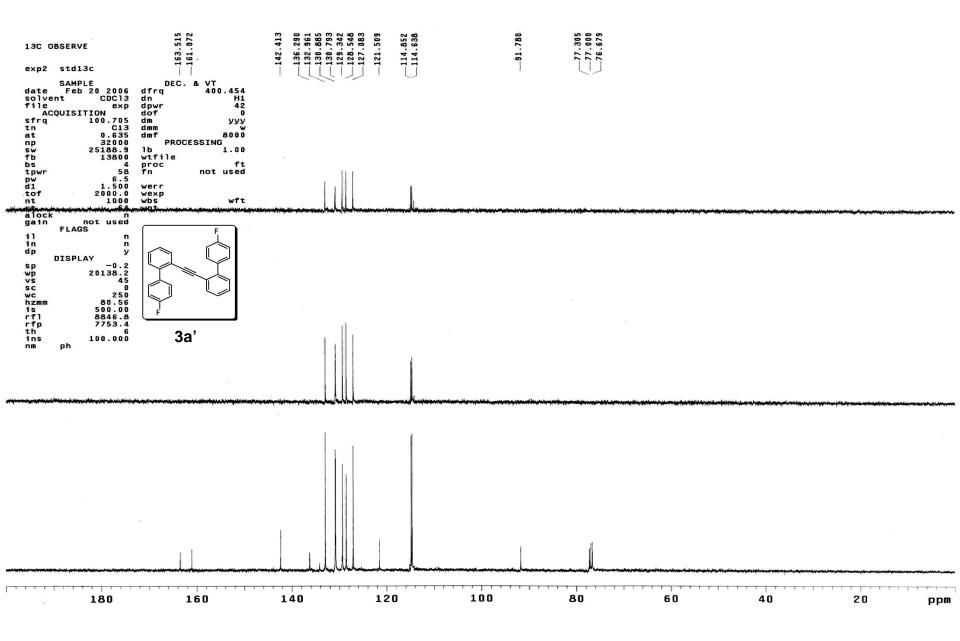


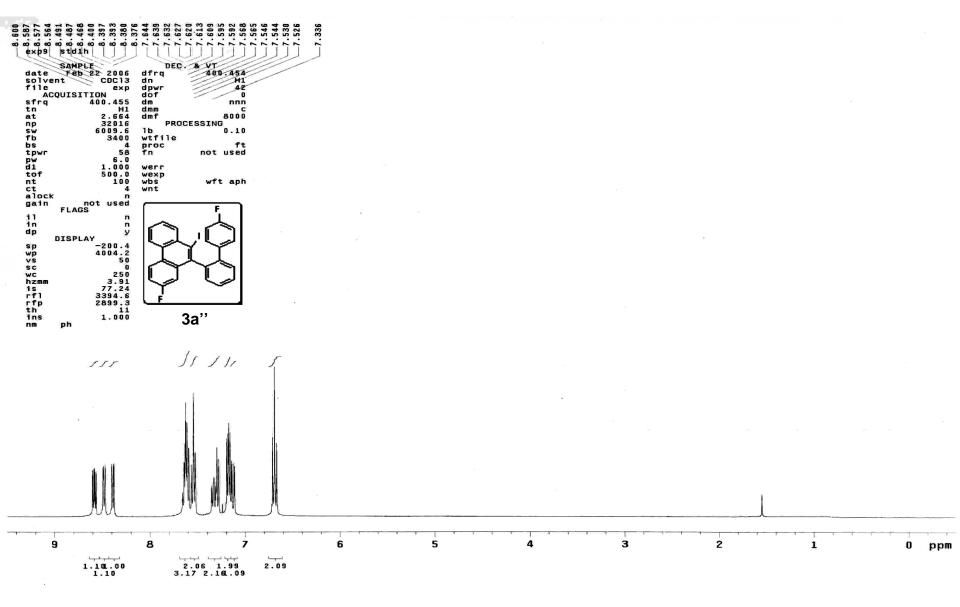


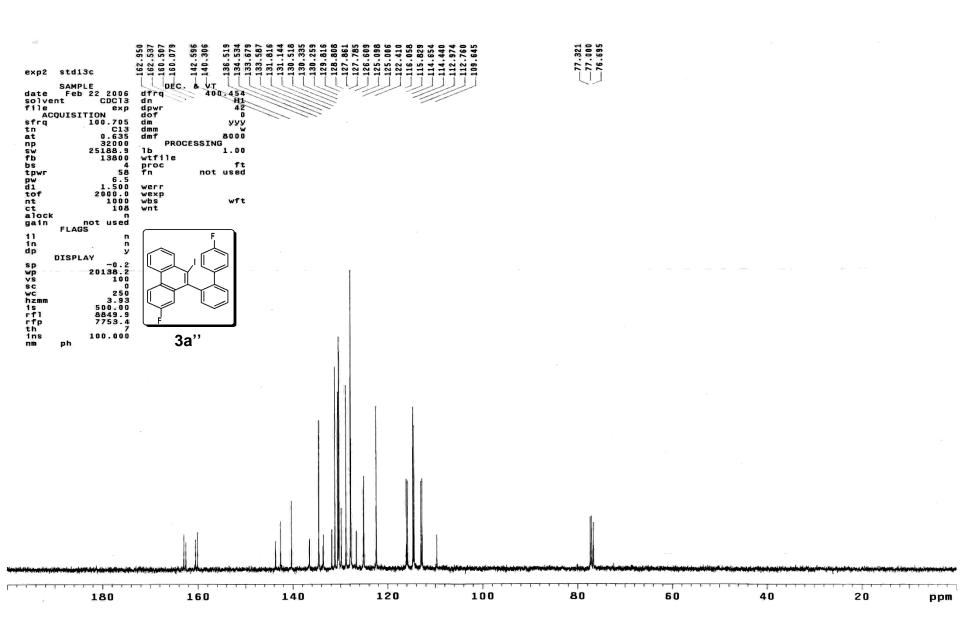


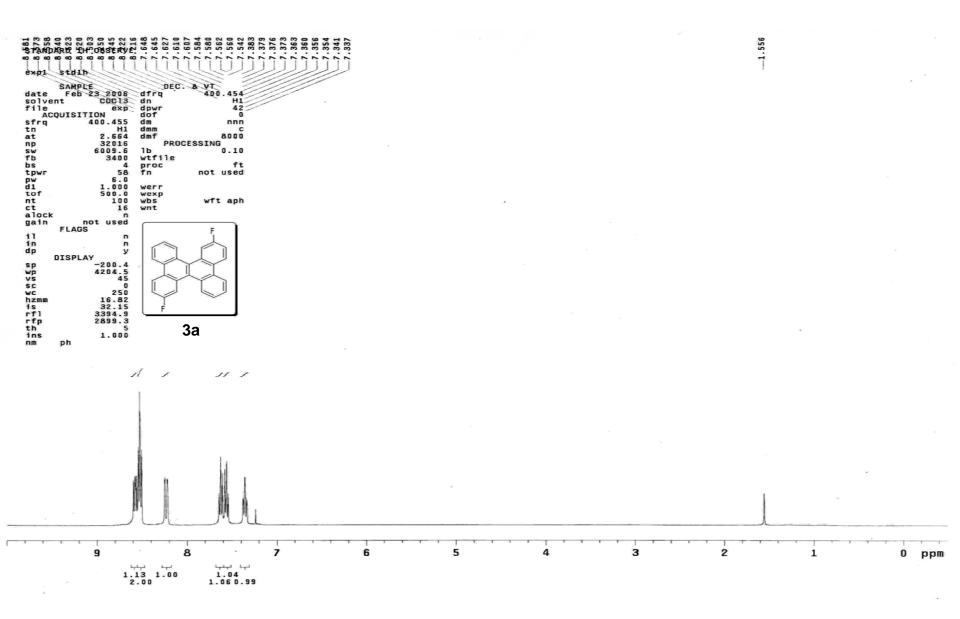


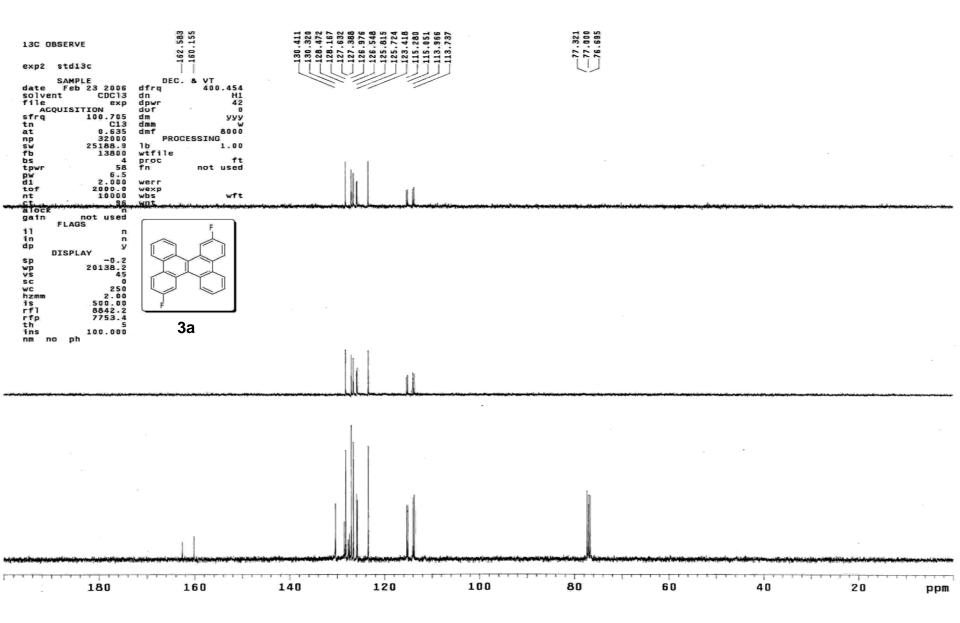


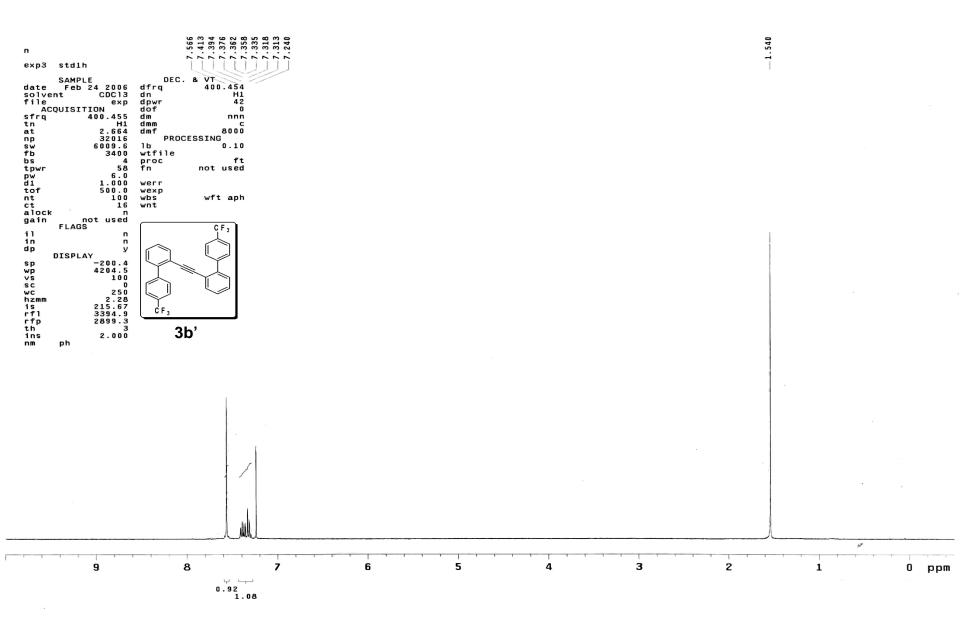


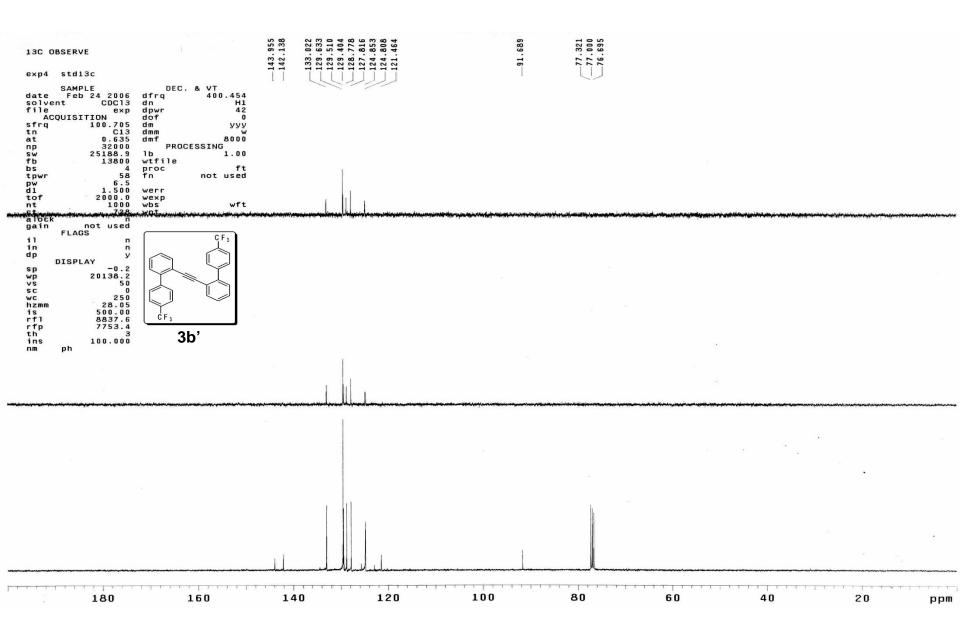




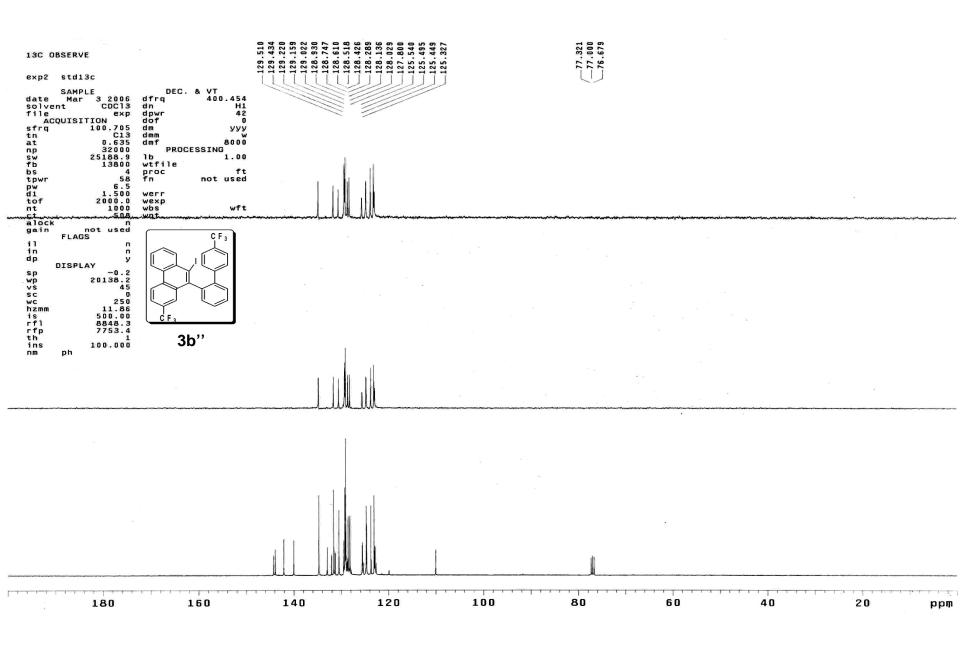


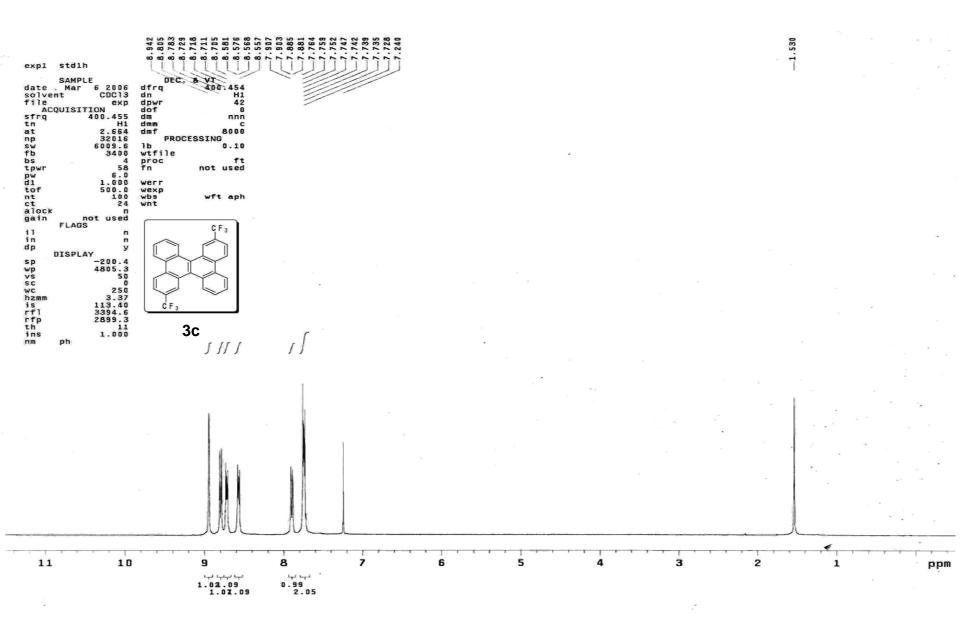


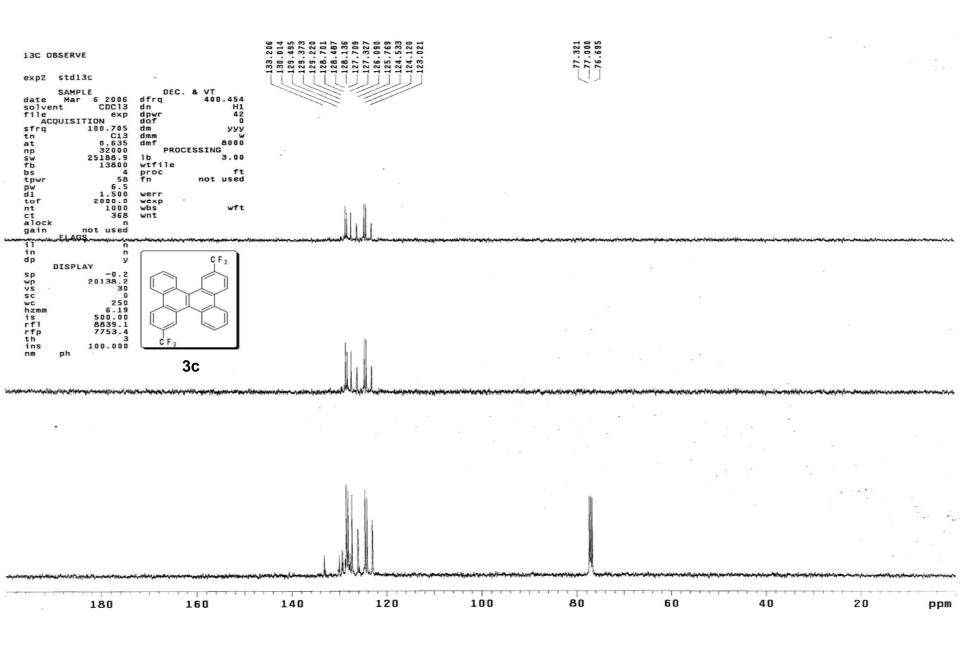


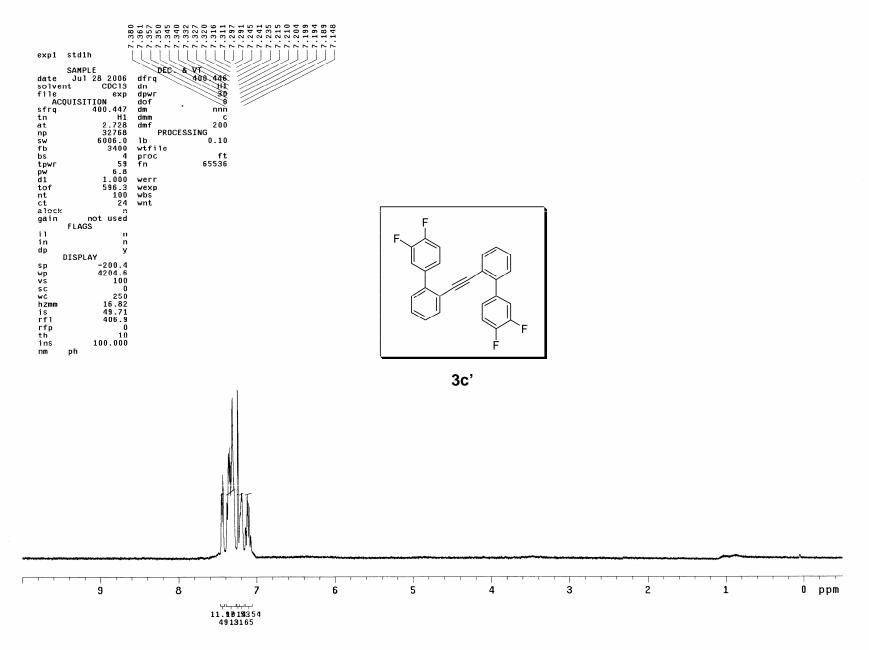


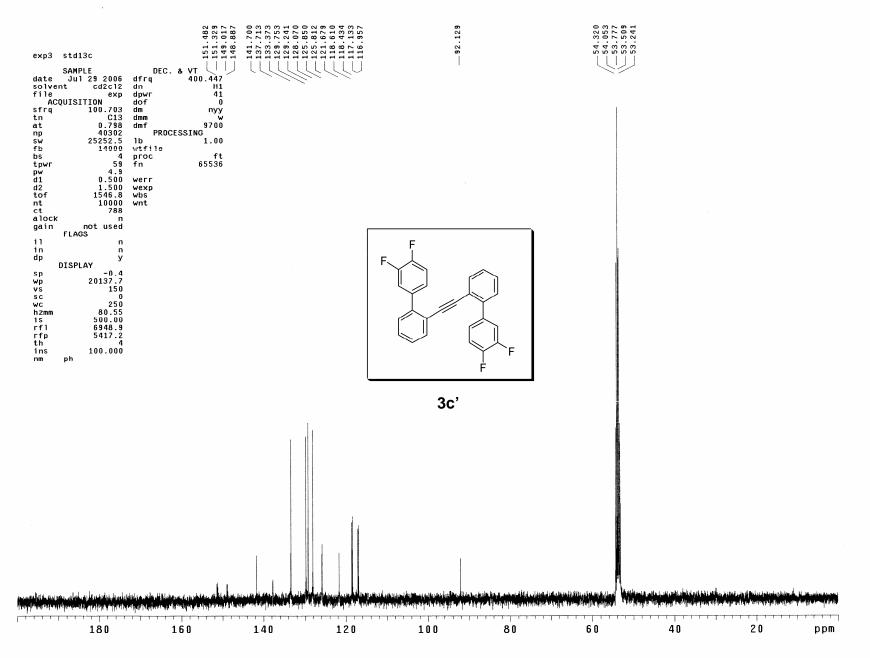
6xb4 2td1 6xb4 2td1 7 7 7 7 6 8 7 7 7 7 8 8		1.541
SAMPLE DEC. & VT date Mar 1 2006 dfrq 400.454 solvent CDC13 dn H1 file exp dpwr 42 ACQUISITION dof 0 sfrq 400.455 dm nnn tn H1 dmm c at 2.664 dmf 8000 np 32016 PROCESSING		
sw 6009.6 b 0.10 fb 3400 wtfile bs 4 proc ft tpwr 58 fn not used pw 6.0 - - d1 1.000 werr - - tof 500.0 wexp - - nt 100 wbs wft aph -		
alock gain not used FLAGS fLAGS fLAGS fLAGS fLAGS fF FF		
hzmm 2.73 1s 125.00 rf1 495.3 rfp 0 th 8 3b'' ins 1.000 nm ph		•
		0-
9 8 7 6 5 	5 4 3	2 1 0 ppm

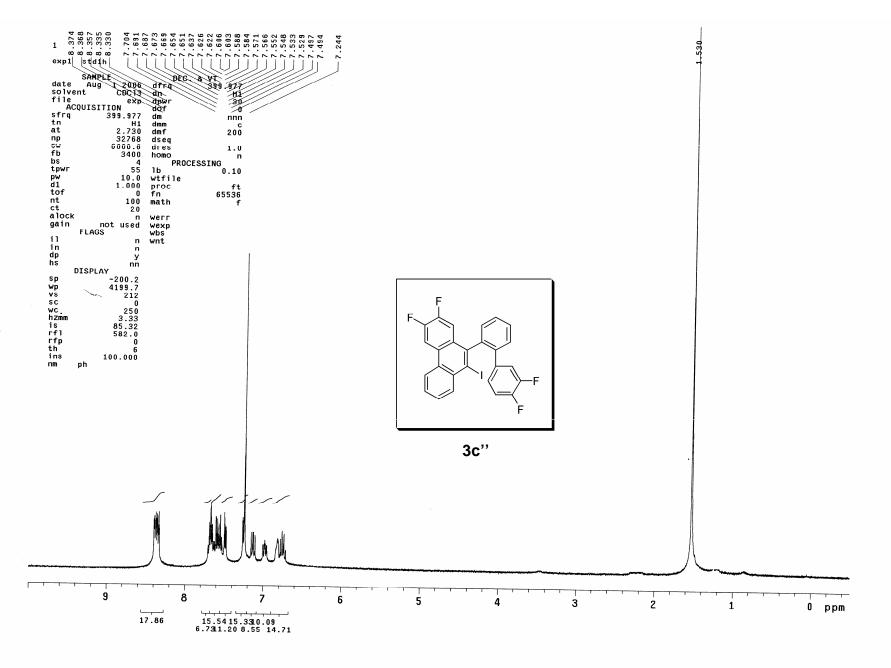


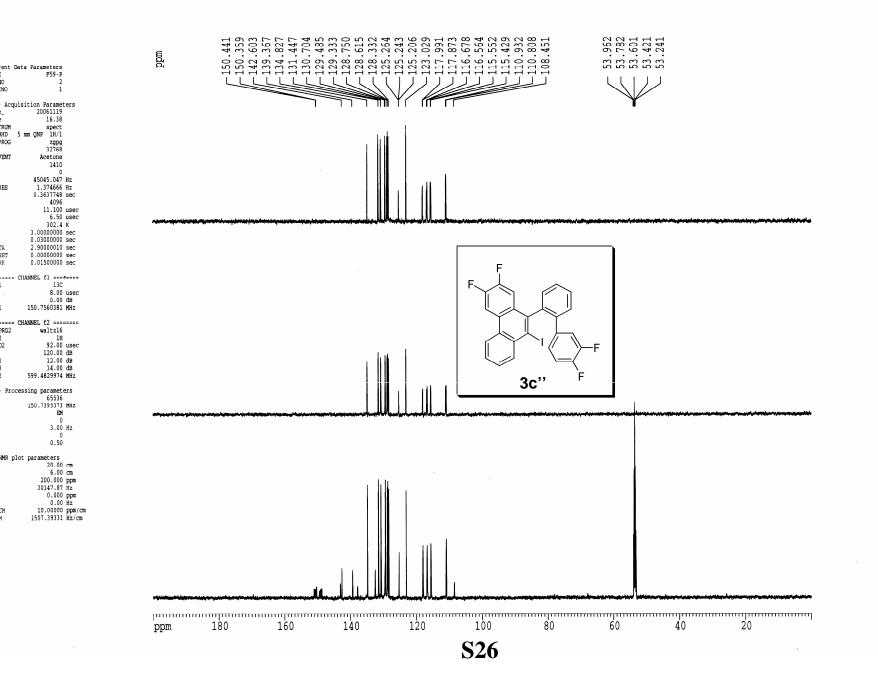


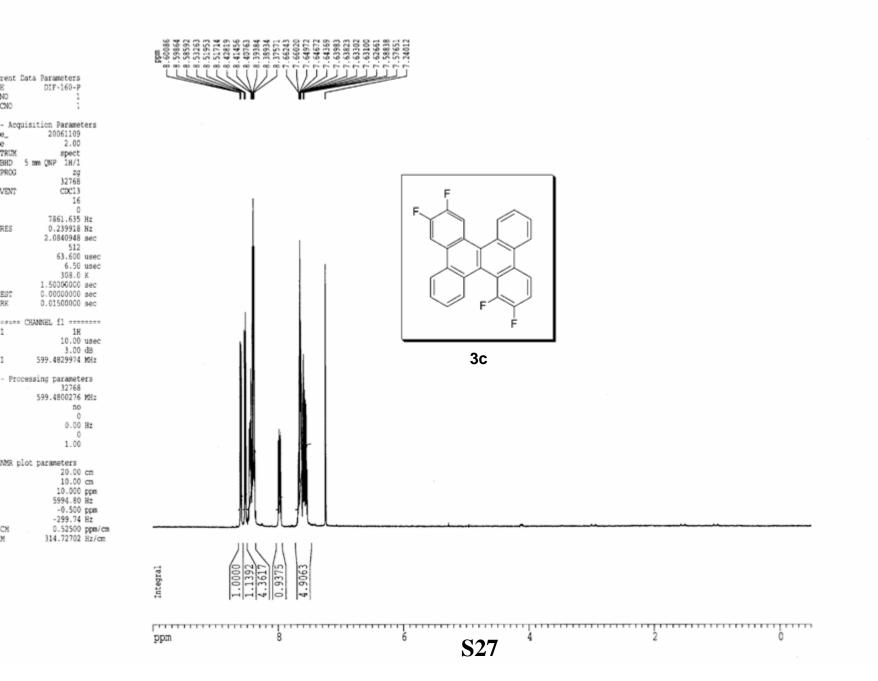


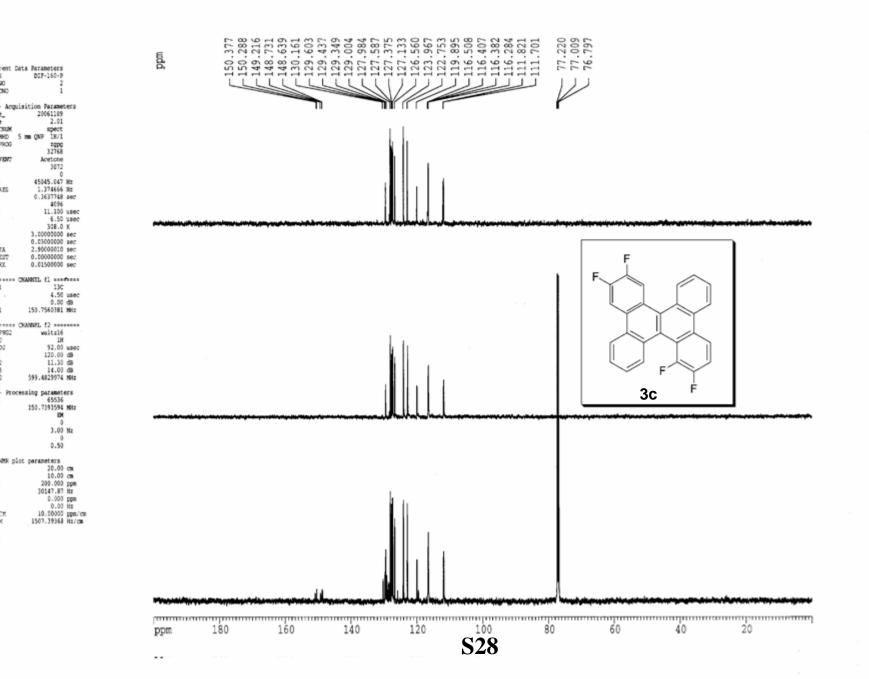


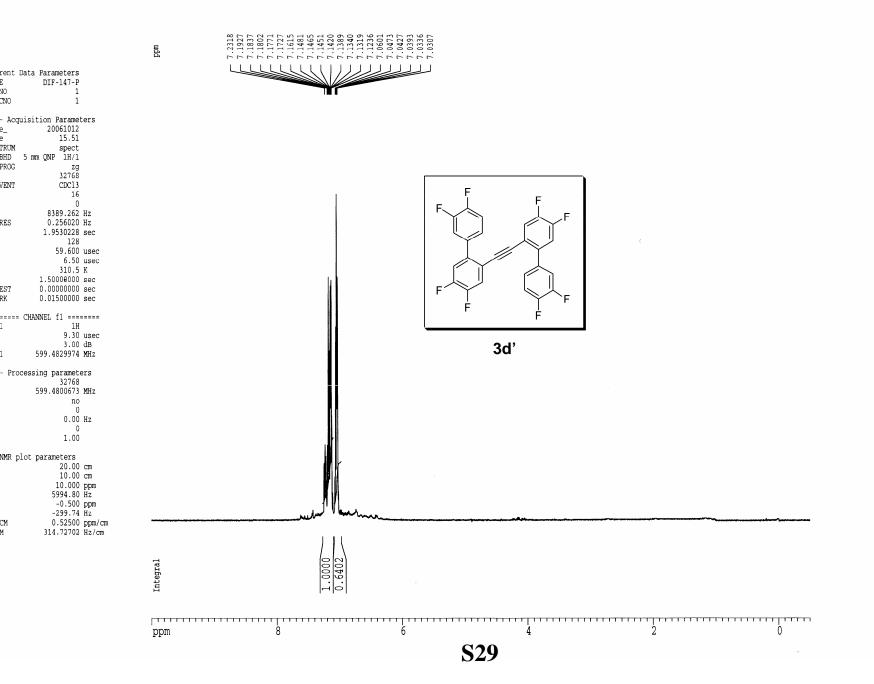


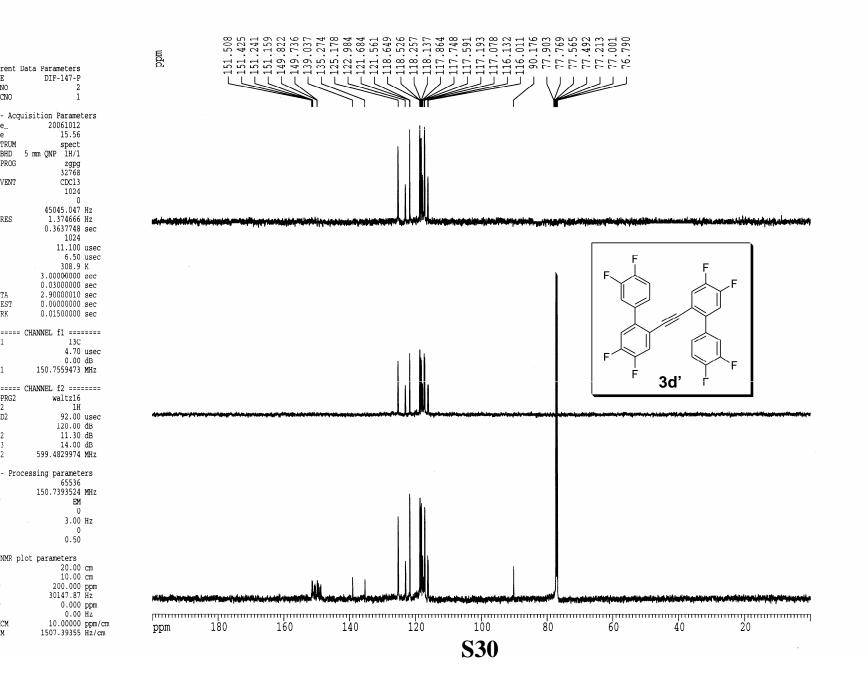


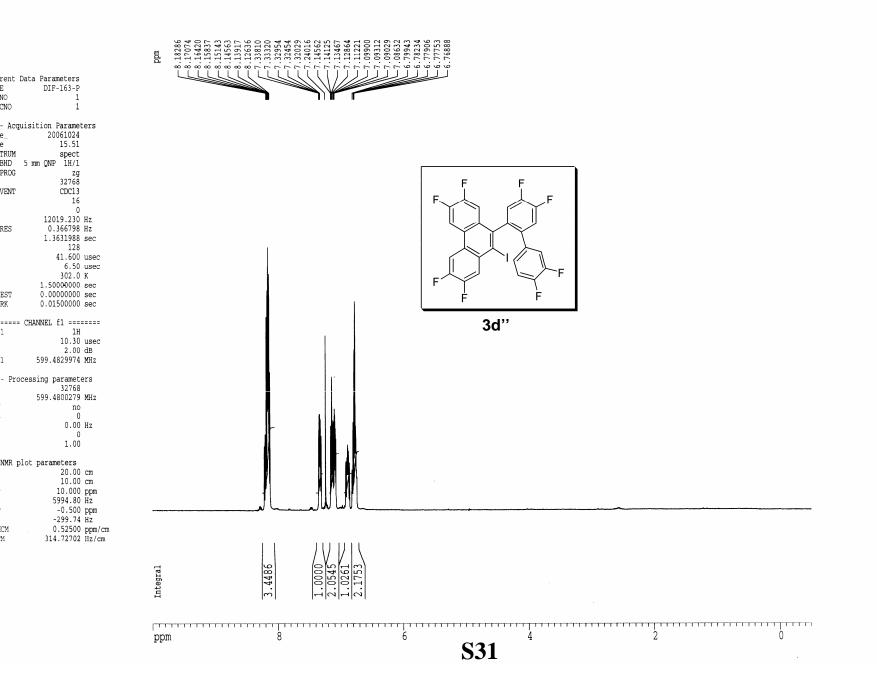


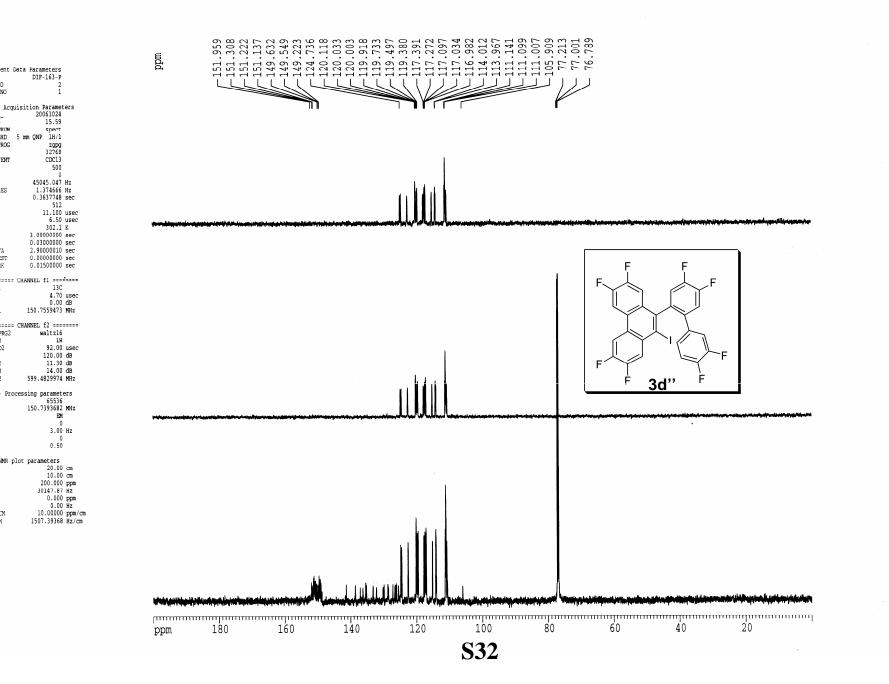


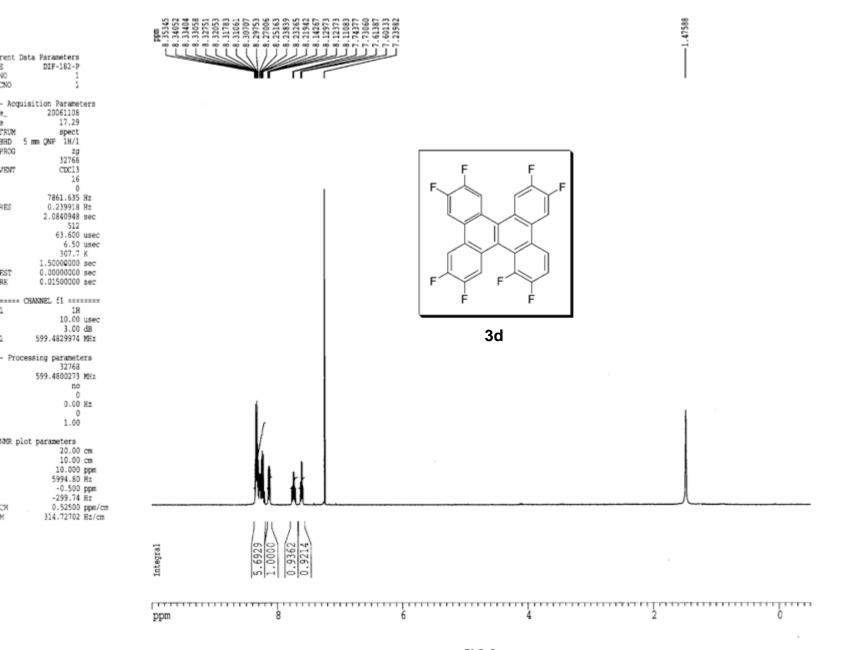


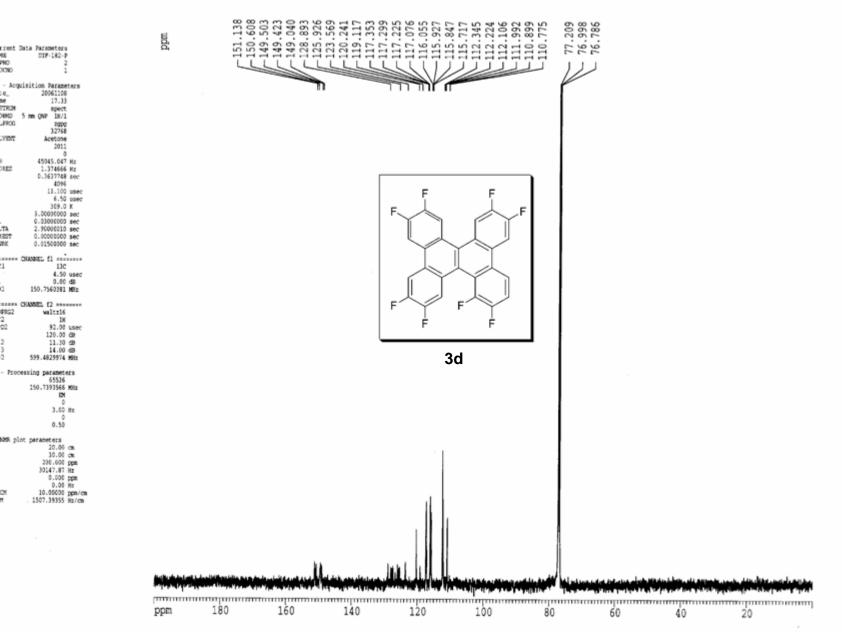


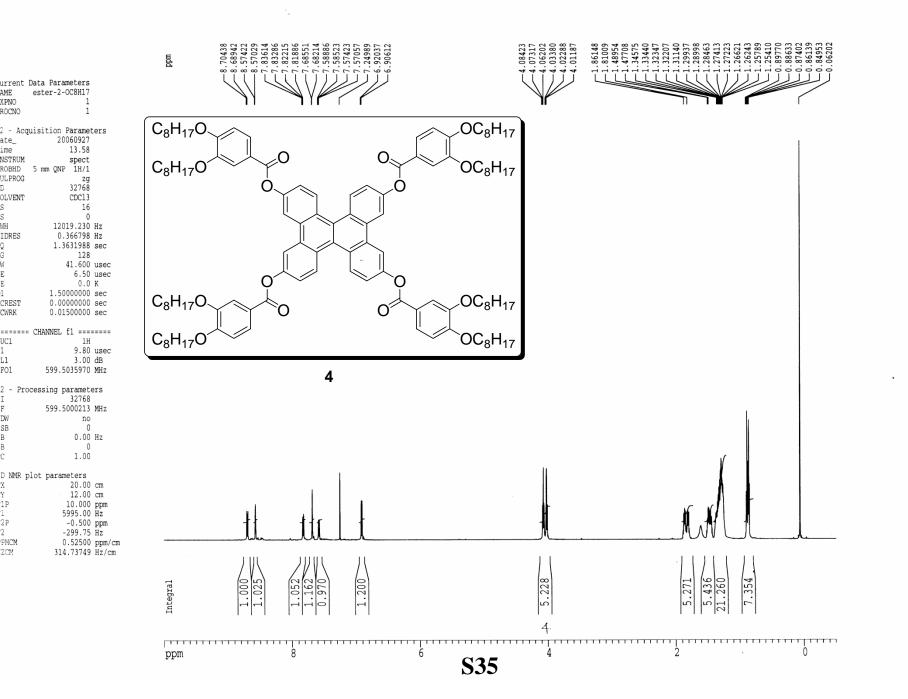


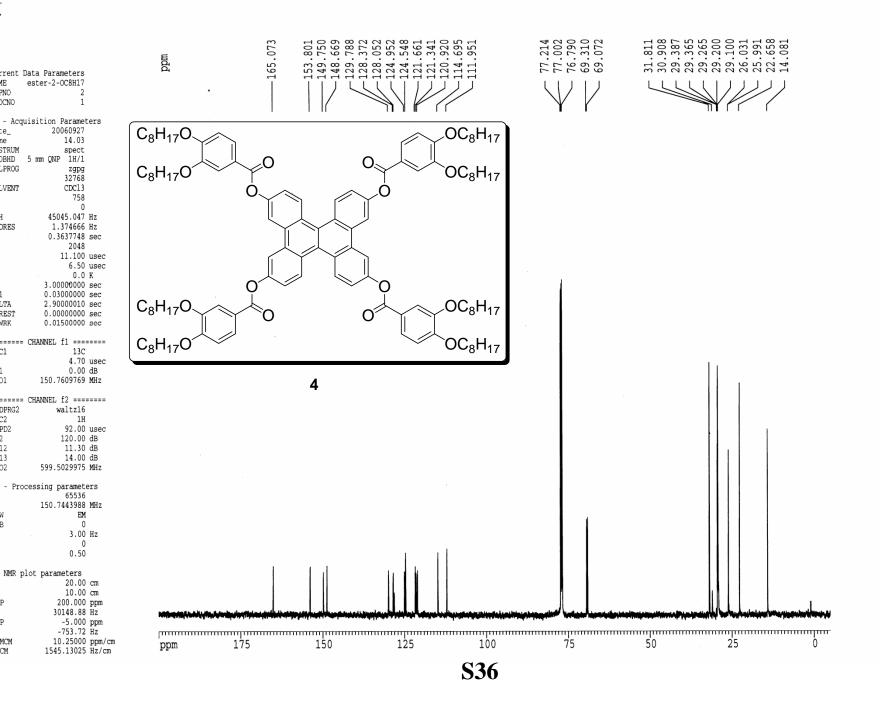












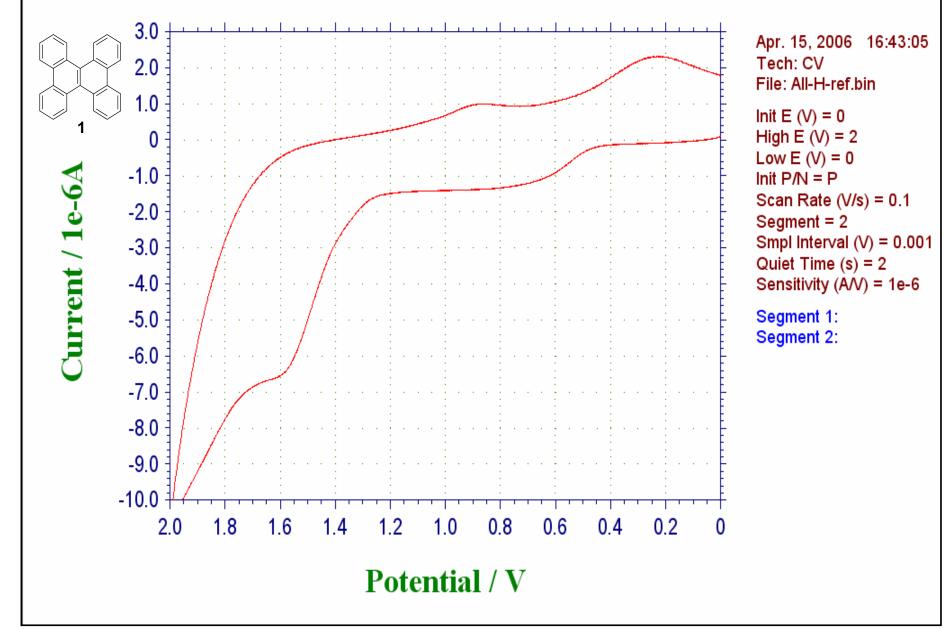


Figure S1. Cyclic voltammetry of 1: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as reference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s)

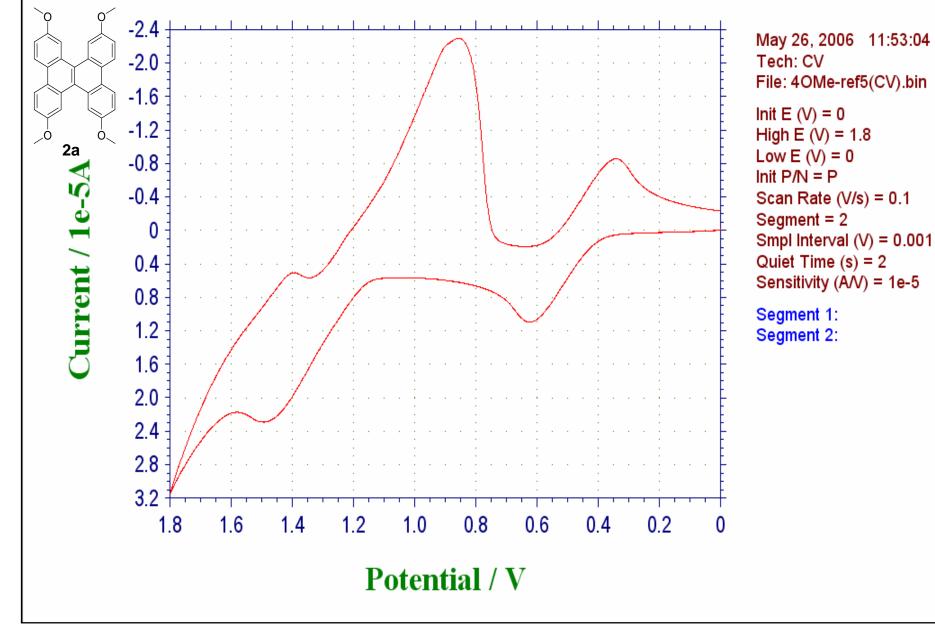


Figure S2. Cyclic voltammetry of **2a**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as reference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s) **S38**

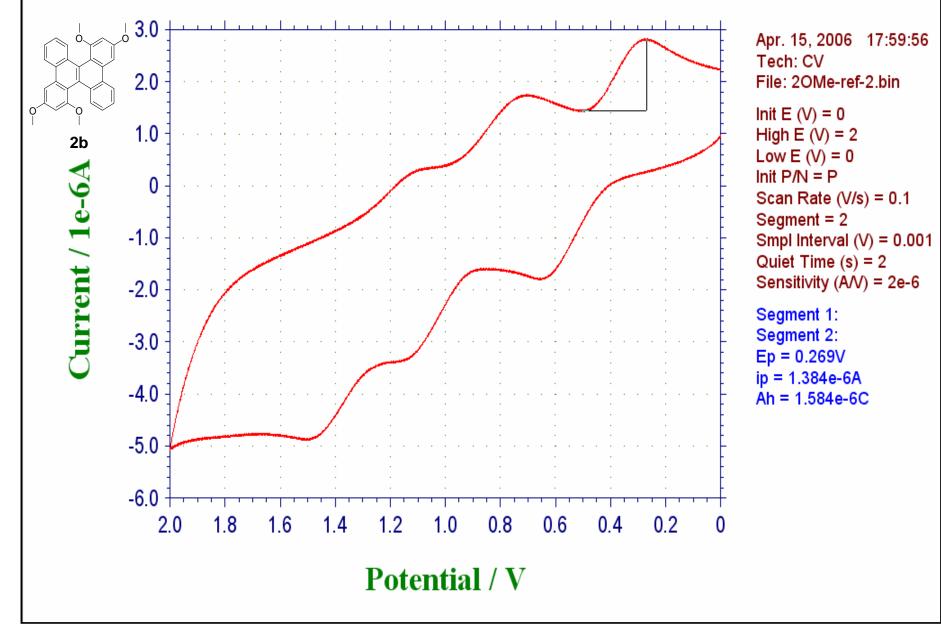


Figure S3. Cyclic voltammetry of **2b**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as reference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s) **S39**

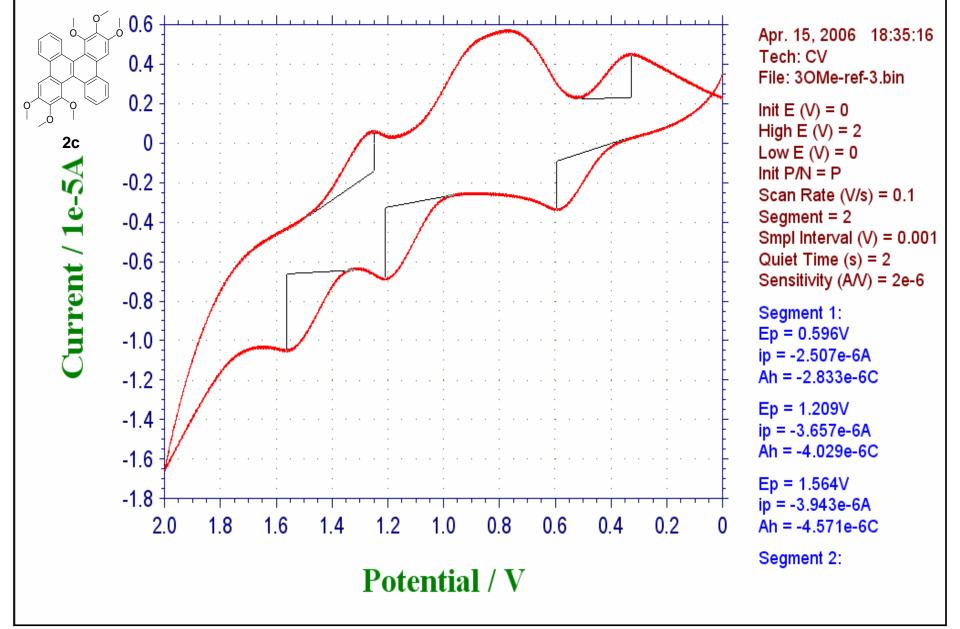


Figure S4. Cyclic voltammetry of **2c**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as reference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s) **S40**

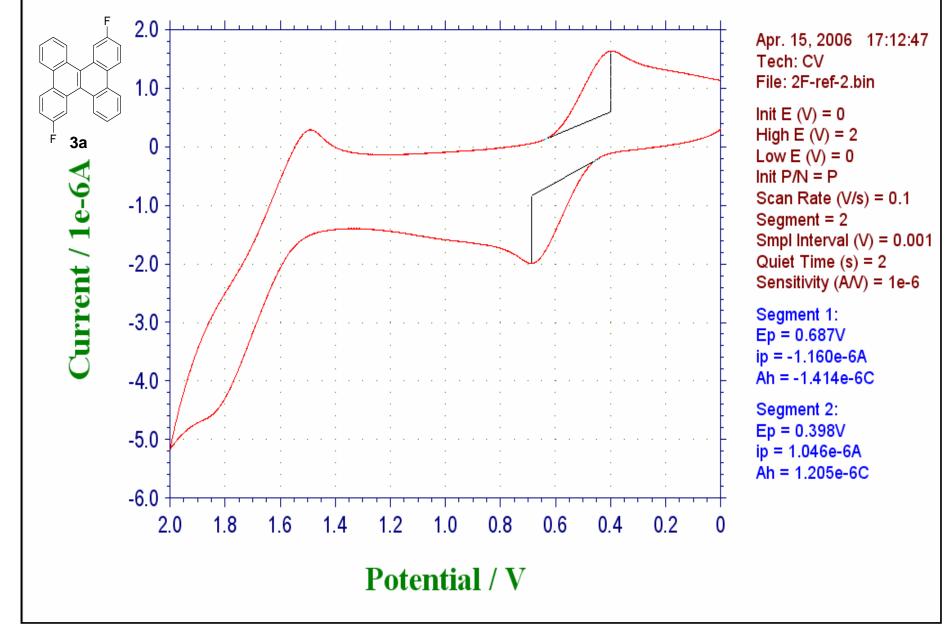


Figure S5. Cyclic voltammetry of **3a**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as reference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s) **S41**

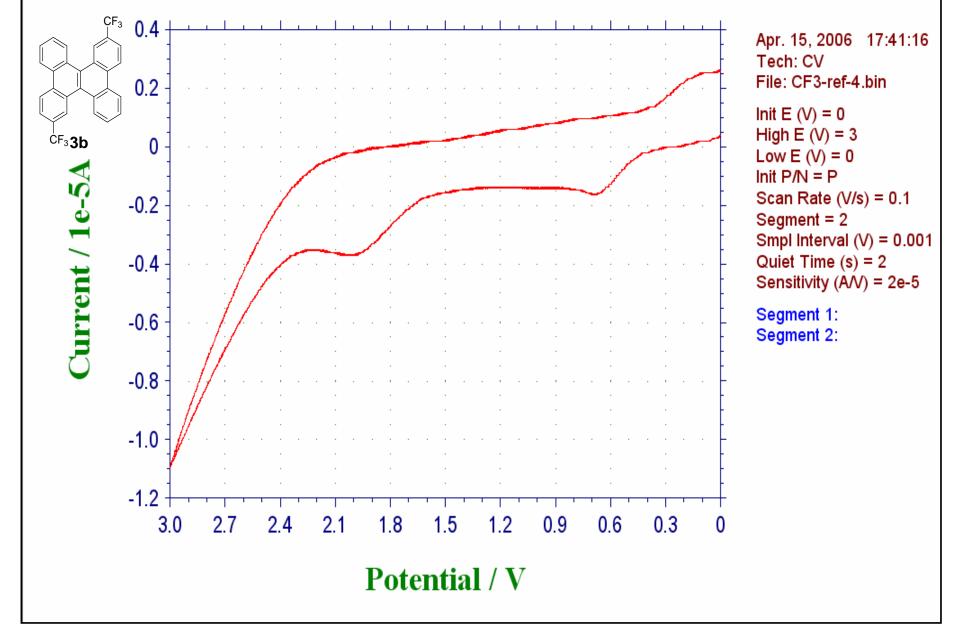


Figure S6. Cyclic voltammetry of **3b**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as efference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s) **S42**

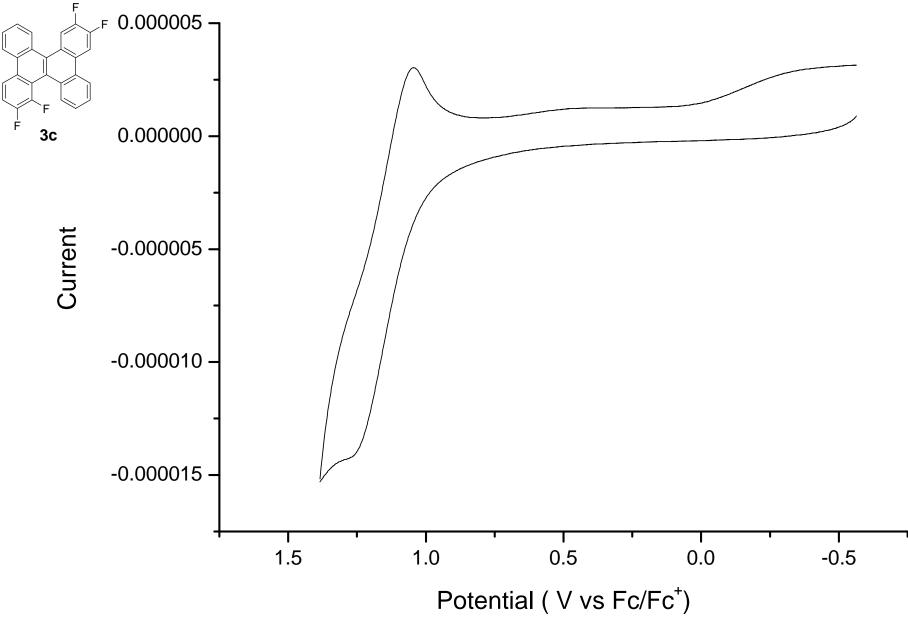


Figure S7. Cyclic voltammetry of **3c**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as efference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s)

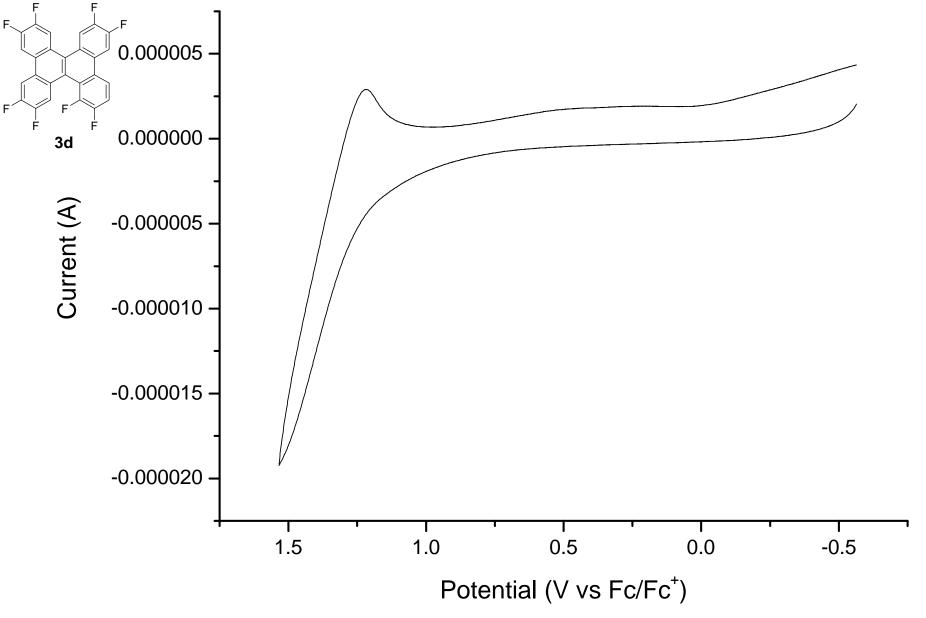


Figure S8. Cyclic voltammetry of **3d**: (Sample concentration was $1*10^{-3}$ M in CH₂Cl₂ for oxidation, with Ag/AgCl as efference electrolyte, Pt as the support electrolyte and glassy carbon electrode as working electrode. The scan rate was 0.1 V/s)

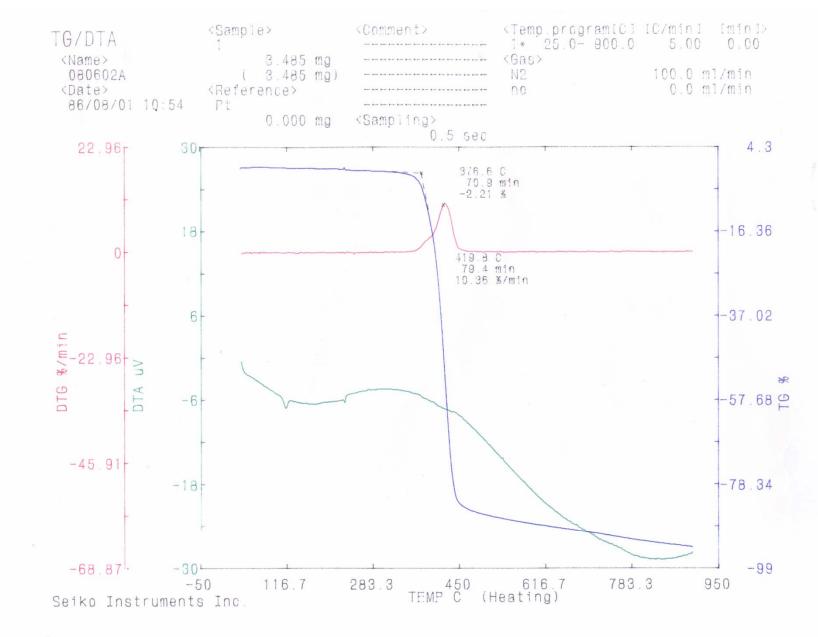


Figure S8. The TGA measurement of 4